

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Ionic Naming Practice

Name the following compounds.

- |                       |                                    |                       |                                       |
|-----------------------|------------------------------------|-----------------------|---------------------------------------|
| 1. KCl                | 2. BaF <sub>2</sub>                | 3. AlBr <sub>3</sub>  | 19. FeCl <sub>2</sub>                 |
| 4. Rb <sub>2</sub> S  | 5. MgO                             | 6. ZnI <sub>2</sub>   | 20. Pb(OH) <sub>2</sub>               |
| 7. SrSe               | 8. Cd <sub>3</sub> N <sub>2</sub>  | 9. Ag <sub>2</sub> S  | 21. NH <sub>4</sub> NO <sub>2</sub>   |
| 10. CsI               | 11. CaTe                           | 12. SrF <sub>2</sub>  | 22. CuI                               |
| 13. Li <sub>3</sub> N | 14. Rb <sub>3</sub> P              | 15. K <sub>2</sub> Se | 23. BaHPO <sub>4</sub>                |
| 16. MgTe              | 17. Mg <sub>3</sub> N <sub>2</sub> | 18. SrCl <sub>2</sub> | 24. Cr(PO <sub>3</sub> ) <sub>2</sub> |

Write formulas for the following compounds.

- |                         |                      |                              |
|-------------------------|----------------------|------------------------------|
| 1. sodium iodide        | 2. Calcium iodide    | 3. Calcium hydroxide         |
| 4. Potassium nitride    | 5. Silver sulfate    | 6. Sodium hydrogen carbonate |
| 7. Calcium phosphate    | 8. Sodium nitrate    | 9. Calcium sulfide           |
| 10. Zinc oxide          | 11. Aluminum cyanide | 12. Potassium bromide        |
| 13. Potassium carbonate | 14. Cadmium nitrate  | 15. Sodium hydroxide         |
| 16. Cesium sulfate      | 17. Lithium selenide | 18. Ammonium chloride        |
| 19. iron(II) hydroxide  | 20. silver acetate   | 21. rubidium arsenide        |